

## Ethyl 3-[2-benzyl-1-(phenylsulfonyl)-1*H*-indol-3-yl]acrylate

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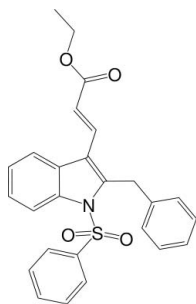
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Key indicators: single-crystal X-ray study; *T* = 100 K; mean  $\sigma(\text{C}-\text{C})$  = 0.002 Å; *R* factor = 0.048; *wR* factor = 0.119; data-to-parameter ratio = 27.4.

In the title compound, C<sub>26</sub>H<sub>23</sub>NO<sub>4</sub>S, the planes of the ethyl acrylate group, the sulfonyl-bound phenyl ring and the benzyl phenyl group are inclined at angles of 7.91 (4), 81.39 (4) and 88.33 (4)°, respectively, with respect to the indole ring system. The conformations of the phenylsulfonyl and benzyl substituents with respect to the indole ring system are influenced by intramolecular C—H···O and C—H··· $\pi$  interactions. Inversion-related molecules at (*x*, *y*, *z*) and (−*x*, 1 − *y*, −*z*) are linked into a centrosymmetric dimer by C—H···O hydrogen bonds. The dimers are crosslinked through C—H··· $\pi$  interactions, forming a two-dimensional network parallel to the *bc* plane. Weak  $\pi$ – $\pi$  stacking interactions are present between the phenyl rings of the benzyl groups; the centroid-to-centroid distance between the rings is 3.8027 (9) Å and the perpendicular distance is 3.743 Å.

### Related literature

For general background and related structures, see: Senthil Kumar *et al.* (2006*a,b,c,d*).



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### Experimental

#### Crystal data

C<sub>26</sub>H<sub>23</sub>NO<sub>4</sub>S  
*M<sub>r</sub>* = 445.51  
 Monoclinic, *P*2<sub>1</sub>/*c*  
*a* = 11.4683 (2) Å  
*b* = 8.8983 (1) Å  
*c* = 21.5088 (3) Å  
 $\beta$  = 99.406 (1)°  
*V* = 2165.43 (5) Å<sup>3</sup>  
*Z* = 4  
 Mo *K*α radiation  
 $\mu$  = 0.18 mm<sup>−1</sup>  
*T* = 100.0 (1) K  
 0.28 × 0.25 × 0.24 mm

#### Data collection

Bruker SMART APEX2 CCD area-detector diffractometer  
 Absorption correction: multi-scan (*SADABS*; Bruker, 2005)  
*T<sub>min</sub>* = 0.912, *T<sub>max</sub>* = 0.958  
 47728 measured reflections  
 7943 independent reflections  
 5869 reflections with *I* > 2σ(*I*)  
*R<sub>int</sub>* = 0.060

#### Refinement

*R*[*F*<sup>2</sup> > 2σ(*F*<sup>2</sup>)] = 0.048  
*wR*(*F*<sup>2</sup>) = 0.119  
*S* = 1.06  
 7943 reflections  
 290 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}}$  = 0.41 e Å<sup>−3</sup>  
 $\Delta\rho_{\text{min}}$  = −0.54 e Å<sup>−3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

<i>D</i> —H··· <i>A</i>	<i>D</i> —H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> —H··· <i>A</i>
C7—H7···O2	0.95	2.36	2.9424 (17)	119
C15—H15A···O1	0.99	2.24	2.8474 (17)	118
C18—H18···O3 <sup>i</sup>	0.95	2.59	3.282 (2)	130
C10—H10···Cg1	0.95	2.75	3.3524 (14)	122
C11—H11···Cg2 <sup>ii</sup>	0.95	2.81	3.6263 (15)	145

Symmetry codes: (i) −*x*, −*y* + 1, −*z*; (ii) −*x*, *y* − ½, −*z* + ½.

Data collection: *APEX2* (Bruker, 2005); cell refinement: *APEX2*; data reduction: *SAINT* (Bruker, 2005); program(s) used to solve structure: *SHELXTL* (Sheldrick, 1998); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL* and *PLATON* (Spek, 2003).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WN2193).

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**supplementary materials**

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## Ethyl 3-[2-benzyl-1-(phenylsulfonyl)-1*H*-indol-3-yl]acrylate

K. Chinnakali, R. Surendran, R. Balamurugan, A. K. Mohanakrishnan and H.-K. Fun

### Comment

The background to this study is set out in Senthil Kumar *et al.* (2006*a,b,c,d*). We now describe the X-ray crystal structure determination of the title compound.

Bond distances and angles in the title compound are comparable to those observed for ethyl 3-(2-methoxycarbonylmethyl-1-phenylsulfonyl-1*H*-indol-3-yl)acrylate (Senthil Kumar *et al.*, 2006*c*). The indole ring system is planar with atom C5 deviating by a maximum of 0.027 (1) Å, from the mean plane through that ring system. As observed in other phenylsulfonylindoles (Senthil Kumar *et al.*, 2006*a,b,c,d*), atom S1 has a distorted tetrahedral configuration, with angles O1—S1—O2 [120.60 (6)°] and N1—S1—C9 [104.87 (6)°] deviating significantly from ideal tetrahedral values. The conformation of the phenylsulfonyl group with respect to the indole unit is described by the torsion angles O1—S1—N1—C1 = 22.55 (12)°, O2—S1—N1—C8 = -36.63 (11)° and N1—S1—C9—C10 = 98.91 (11)°. This conformation is influenced by the intramolecular C—H...O interactions, C7—H7...O2 and C15—H15A...O1, involving the sulfonyl O1 and O2 atoms which deviate 0.143 (1) and 0.509 (1) Å, respectively, from the plane of the indole ring system.

The dihedral angle between the C9—C14 phenyl ring and indole ring system is 81.39 (4)°. The torsion angle N1—C1—C15—C16 of 81.74 (15)° describes the conformation of the attachment of the benzyl substituent to the indole ring system and torsion angle C1—C15—C16—C21 of 24.17 (18)° shows how the C16—C21 phenyl ring is oriented. This conformation is influenced by the intramolecular C15—H15A...O1 interaction, and by the C—H... $\pi$  interaction involving H10 and the C16—C21 ring, with H10 separated from the ring centroid (*Cg*1) by 2.75 Å (Table 1). The dihedral angle between the mean planes through the C9—C14 and C16—C21 aromatic rings is 33.35 (6)°; the centroids of these two rings are separated by 4.4737 (8) Å and hence there is no  $\pi$ - $\pi$  interaction between them.

Atoms O3, O4, C22 to C26 of the ethyl acrylate substituent group at C2 are coplanar (r.m.s deviation of 0.054 Å). This plane is slightly twisted away from the indole ring system by an angle of 7.91 (4)°.

The C18—H18...O3 hydrogen bonds link the inversion-related molecules at (*x*, *y*, *z*) and (-*x*, 1 - *y*, -*z*) into a centrosymmetric dimer. The dimer structure is further stabilized by a weak  $\pi$ - $\pi$  stacking interaction between the phenyl rings (C16—C21) of the benzyl groups; the centroid-centroid distance between the rings is 3.8027 (9) Å and the perpendicular distance is 3.743 Å. The dimers are linked through C—H... $\pi$  interactions (Table 1) involving H11 and the benzene ring (C3—C8; centroid *Cg*2) of the indole ring system, to form a two-dimensional network parallel to the *bc* plane (Fig. 2).

### Experimental

(*E*)-Ethyl 3-[2-(hydroxymethyl)-1-(phenylsulfonyl)-1*H*-indol-3-yl]acrylate (1.3 mmol) and boron trifluoride etherate solution (5 drops) in dry benzene (20 ml) were refluxed for 4 h. After completion of the reaction, the mixture was poured into ice-water, washed with water (3 × 10 ml) followed by solvent removal to give the crude product. It was recrystallized from methanol (5 ml).

## Refinement

H atoms were positioned geometrically and were treated as riding on their parent C atoms, with C—H distances of 0.95 Å ( $C_{sp^2}$ ), 0.98 Å (methyl) and 0.99 Å (methylene), and  $U_{iso}(H)$  value of  $1.2U_{eq}(C)$  or  $1.5U_{eq}(C_{methyl})$

## Figures

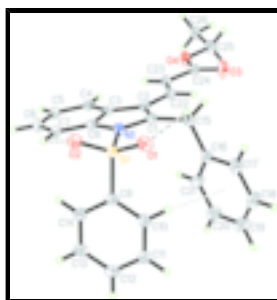


Fig. 1. Molecular structure of the title compound, showing the atomic numbering and 50% probability displacement ellipsoids. C—H...O hydrogen bonds are shown as dashed lines and the dotted line represents a C—H... $\pi$  interaction.

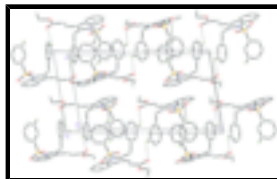


Fig. 2. Part of the crystal packing of the title compound, viewed approximately down the  $b$  axis. C—H...O hydrogen bonds are shown as dashed lines, C—H... $\pi$  interactions are shown as double dashed lines and  $\pi$ - $\pi$  interactions are shown as dotted lines. H atoms not involved in these interactions have been omitted.

## Ethyl 3-[2-benzyl-1-(phenylsulfonyl)-1*H*-indol-3-yl]acrylate

### Crystal data

$C_{26}H_{23}NO_4S$

$M_r = 445.51$

Monoclinic,  $P2_1/c$

Hall symbol: -P 2ybc

$a = 11.4683$  (2) Å

$b = 8.8983$  (1) Å

$c = 21.5088$  (3) Å

$\beta = 99.406$  (1)°

$V = 2165.43$  (5) Å<sup>3</sup>

$Z = 4$

$F_{000} = 936$

$D_x = 1.367$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation

$\lambda = 0.71073$  Å

Cell parameters from 7382 reflections

$\theta = 2.4$ – $30.6$ °

$\mu = 0.18$  mm<sup>-1</sup>

$T = 100.0$  (1) K

Block, yellow

$0.28 \times 0.25 \times 0.24$  mm

### Data collection

Bruker SMART APEX2 CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

Detector resolution: 8.33 pixels mm<sup>-1</sup>

$T = 100.0$ (1) K

7943 independent reflections

5869 reflections with  $I > 2\sigma(I)$

$R_{int} = 0.060$

$\theta_{max} = 32.7$ °

$\theta_{min} = 1.9$ °

$\omega$  scans  $h = -17 \rightarrow 14$   
 Absorption correction: multi-scan  $k = -12 \rightarrow 13$   
 (SADABS; Bruker, 2005)  
 $T_{\min} = 0.912$ ,  $T_{\max} = 0.958$   $l = -32 \rightarrow 31$   
 47728 measured reflections

### Refinement

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.048$	H-atom parameters constrained
$wR(F^2) = 0.119$	$w = 1/[\sigma^2(F_o^2) + (0.0462P)^2 + 0.737P]$
$S = 1.06$	where $P = (F_o^2 + 2F_c^2)/3$
7943 reflections	$(\Delta/\sigma)_{\max} = 0.001$
290 parameters	$\Delta\rho_{\max} = 0.41 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\min} = -0.54 \text{ e } \text{\AA}^{-3}$
	Extinction correction: none

### Special details

**Experimental.** The data was collected with the Oxford Cyrosystem Cobra low-temperature attachment.

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > 2\sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.27475 (3)	0.34607 (4)	0.252973 (15)	0.01888 (8)
O1	0.26409 (9)	0.23974 (11)	0.20278 (5)	0.0246 (2)
O2	0.35867 (8)	0.32365 (11)	0.30905 (5)	0.0235 (2)
O3	0.29725 (11)	0.91913 (13)	-0.02392 (5)	0.0337 (3)
O4	0.37486 (9)	1.13030 (11)	0.02295 (4)	0.0255 (2)
N1	0.31080 (9)	0.51204 (12)	0.22307 (5)	0.0177 (2)
C1	0.29109 (11)	0.55791 (14)	0.15938 (6)	0.0173 (2)
C2	0.32209 (11)	0.70722 (14)	0.15648 (6)	0.0168 (2)
C3	0.36031 (10)	0.75975 (14)	0.22039 (6)	0.0164 (2)
C4	0.39607 (11)	0.90026 (15)	0.24600 (6)	0.0181 (2)
H4	0.4010	0.9843	0.2193	0.022*
C5	0.42414 (12)	0.91474 (16)	0.31074 (6)	0.0214 (3)

## supplementary materials

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H5	0.4468	1.0101	0.3286	0.026*
C6	0.41952 (12)	0.79078 (16)	0.35020 (6)	0.0223 (3)
H6	0.4410	0.8029	0.3945	0.027*
C7	0.38426 (11)	0.65075 (15)	0.32612 (6)	0.0203 (2)
H7	0.3812	0.5667	0.3530	0.024*
C8	0.35341 (11)	0.63749 (14)	0.26097 (6)	0.0169 (2)
C9	0.13509 (11)	0.37614 (14)	0.27434 (6)	0.0190 (2)
C10	0.03426 (12)	0.32885 (16)	0.23390 (6)	0.0231 (3)
H10	0.0407	0.2792	0.1955	0.028*
C11	-0.07581 (12)	0.35557 (17)	0.25069 (7)	0.0261 (3)
H11	-0.1455	0.3244	0.2236	0.031*
C12	-0.08422 (12)	0.42756 (16)	0.30685 (7)	0.0253 (3)
H12	-0.1598	0.4472	0.3177	0.030*
C13	0.01697 (12)	0.47128 (16)	0.34750 (7)	0.0250 (3)
H13	0.0102	0.5183	0.3864	0.030*
C14	0.12761 (12)	0.44652 (15)	0.33152 (6)	0.0221 (3)
H14	0.1971	0.4769	0.3590	0.027*
C15	0.24417 (12)	0.45820 (15)	0.10488 (6)	0.0206 (3)
H15A	0.2807	0.3576	0.1124	0.025*
H15B	0.2691	0.4996	0.0664	0.025*
C16	0.11042 (12)	0.43986 (15)	0.09285 (6)	0.0206 (3)
C17	0.06111 (14)	0.31429 (17)	0.06005 (7)	0.0265 (3)
H17	0.1112	0.2417	0.0456	0.032*
C18	-0.06059 (14)	0.29404 (18)	0.04830 (7)	0.0319 (3)
H18	-0.0932	0.2079	0.0258	0.038*
C19	-0.13458 (14)	0.39880 (19)	0.06916 (7)	0.0317 (3)
H19	-0.2177	0.3838	0.0617	0.038*
C20	-0.08669 (13)	0.52580 (19)	0.10099 (7)	0.0305 (3)
H20	-0.1371	0.5985	0.1151	0.037*
C21	0.03539 (13)	0.54664 (16)	0.11225 (7)	0.0253 (3)
H21	0.0677	0.6347	0.1334	0.030*
C22	0.31486 (11)	0.78830 (15)	0.09784 (6)	0.0190 (2)
H22	0.2829	0.7339	0.0610	0.023*
C23	0.34727 (12)	0.93034 (15)	0.08850 (6)	0.0213 (3)
H23	0.3773	0.9928	0.1233	0.026*
C24	0.33564 (12)	0.98804 (15)	0.02363 (6)	0.0208 (3)
C25	0.36841 (14)	1.19570 (17)	-0.03932 (6)	0.0275 (3)
H25A	0.2849	1.2104	-0.0590	0.033*
H25B	0.4067	1.1286	-0.0667	0.033*
C26	0.43137 (16)	1.34387 (17)	-0.03127 (7)	0.0312 (3)
H26A	0.4278	1.3922	-0.0725	0.047*
H26B	0.5141	1.3276	-0.0124	0.047*
H26C	0.3933	1.4087	-0.0037	0.047*

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
S1	0.02132 (15)	0.01453 (14)	0.02155 (15)	0.00192 (11)	0.00570 (12)	0.00244 (11)

O1	0.0317 (5)	0.0159 (4)	0.0278 (5)	0.0009 (4)	0.0100 (4)	-0.0008 (4)
O2	0.0225 (5)	0.0228 (5)	0.0251 (5)	0.0049 (4)	0.0039 (4)	0.0079 (4)
O3	0.0488 (7)	0.0312 (6)	0.0189 (5)	-0.0167 (5)	-0.0013 (4)	0.0008 (4)
O4	0.0393 (6)	0.0207 (5)	0.0161 (4)	-0.0072 (4)	0.0032 (4)	0.0031 (4)
N1	0.0201 (5)	0.0155 (5)	0.0175 (5)	-0.0005 (4)	0.0036 (4)	0.0012 (4)
C1	0.0178 (5)	0.0181 (6)	0.0166 (5)	0.0000 (4)	0.0044 (4)	-0.0004 (4)
C2	0.0165 (5)	0.0170 (5)	0.0171 (5)	-0.0005 (4)	0.0032 (4)	0.0004 (4)
C3	0.0144 (5)	0.0178 (5)	0.0170 (5)	0.0007 (4)	0.0025 (4)	0.0004 (4)
C4	0.0178 (5)	0.0178 (6)	0.0189 (6)	-0.0006 (4)	0.0034 (4)	-0.0006 (5)
C5	0.0213 (6)	0.0222 (6)	0.0200 (6)	-0.0027 (5)	0.0020 (5)	-0.0031 (5)
C6	0.0224 (6)	0.0272 (7)	0.0163 (6)	-0.0009 (5)	0.0000 (5)	-0.0008 (5)
C7	0.0203 (6)	0.0227 (6)	0.0174 (6)	0.0008 (5)	0.0012 (5)	0.0032 (5)
C8	0.0150 (5)	0.0169 (6)	0.0189 (6)	0.0007 (4)	0.0030 (4)	0.0003 (4)
C9	0.0199 (6)	0.0171 (6)	0.0209 (6)	-0.0002 (4)	0.0055 (5)	0.0033 (5)
C10	0.0255 (6)	0.0238 (6)	0.0201 (6)	-0.0054 (5)	0.0044 (5)	0.0004 (5)
C11	0.0211 (6)	0.0310 (7)	0.0258 (7)	-0.0063 (5)	0.0025 (5)	0.0024 (6)
C12	0.0219 (6)	0.0227 (6)	0.0324 (7)	-0.0008 (5)	0.0075 (5)	0.0036 (6)
C13	0.0255 (7)	0.0231 (6)	0.0278 (7)	-0.0008 (5)	0.0087 (5)	-0.0037 (5)
C14	0.0218 (6)	0.0219 (6)	0.0226 (6)	-0.0013 (5)	0.0033 (5)	-0.0015 (5)
C15	0.0258 (6)	0.0176 (6)	0.0186 (6)	-0.0026 (5)	0.0047 (5)	-0.0025 (5)
C16	0.0242 (6)	0.0197 (6)	0.0173 (6)	-0.0038 (5)	0.0015 (5)	0.0015 (5)
C17	0.0327 (7)	0.0235 (7)	0.0222 (6)	-0.0042 (6)	0.0006 (5)	-0.0026 (5)
C18	0.0353 (8)	0.0285 (7)	0.0284 (7)	-0.0109 (6)	-0.0052 (6)	0.0019 (6)
C19	0.0244 (7)	0.0380 (8)	0.0299 (8)	-0.0064 (6)	-0.0037 (6)	0.0095 (7)
C20	0.0252 (7)	0.0334 (8)	0.0317 (8)	0.0019 (6)	0.0009 (6)	0.0039 (6)
C21	0.0256 (7)	0.0227 (7)	0.0269 (7)	-0.0018 (5)	0.0017 (5)	-0.0013 (5)
C22	0.0194 (6)	0.0205 (6)	0.0170 (5)	-0.0018 (5)	0.0025 (4)	0.0001 (5)
C23	0.0261 (6)	0.0214 (6)	0.0157 (6)	-0.0035 (5)	0.0013 (5)	0.0005 (5)
C24	0.0221 (6)	0.0219 (6)	0.0180 (6)	-0.0032 (5)	0.0024 (5)	0.0019 (5)
C25	0.0378 (8)	0.0271 (7)	0.0171 (6)	-0.0044 (6)	0.0030 (5)	0.0067 (5)
C26	0.0492 (9)	0.0227 (7)	0.0239 (7)	-0.0045 (6)	0.0121 (7)	0.0036 (6)

*Geometric parameters (Å, °)*

S1—O1	1.4257 (10)	C12—C13	1.389 (2)
S1—O2	1.4288 (10)	C12—H12	0.95
S1—N1	1.6889 (11)	C13—C14	1.3853 (19)
S1—C9	1.7574 (13)	C13—H13	0.95
O3—C24	1.2118 (16)	C14—H14	0.95
O4—C24	1.3443 (16)	C15—C16	1.5220 (19)
O4—C25	1.4510 (16)	C15—H15A	0.99
N1—C1	1.4118 (16)	C15—H15B	0.99
N1—C8	1.4210 (16)	C16—C21	1.391 (2)
C1—C2	1.3794 (18)	C16—C17	1.3913 (19)
C1—C15	1.4984 (17)	C17—C18	1.389 (2)
C2—C22	1.4436 (18)	C17—H17	0.95
C2—C3	1.4503 (17)	C18—C19	1.384 (2)
C3—C4	1.4004 (18)	C18—H18	0.95
C3—C8	1.4051 (17)	C19—C20	1.387 (2)

## supplementary materials

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C4—C5	1.3830 (18)	C19—H19	0.95
C4—H4	0.95	C20—C21	1.394 (2)
C5—C6	1.3980 (19)	C20—H20	0.95
C5—H5	0.95	C21—H21	0.95
C6—C7	1.3842 (19)	C22—C23	1.3418 (18)
C6—H6	0.95	C22—H22	0.95
C7—C8	1.3930 (18)	C23—C24	1.4722 (18)
C7—H7	0.95	C23—H23	0.95
C9—C10	1.3935 (19)	C25—C26	1.500 (2)
C9—C14	1.3954 (19)	C25—H25A	0.99
C10—C11	1.389 (2)	C25—H25B	0.99
C10—H10	0.95	C26—H26A	0.98
C11—C12	1.384 (2)	C26—H26B	0.98
C11—H11	0.95	C26—H26C	0.98
O1—S1—O2	120.60 (6)	C13—C14—C9	118.74 (12)
O1—S1—N1	106.74 (6)	C13—C14—H14	120.6
O2—S1—N1	105.69 (6)	C9—C14—H14	120.6
O1—S1—C9	109.27 (6)	C1—C15—C16	114.81 (11)
O2—S1—C9	108.51 (6)	C1—C15—H15A	108.6
N1—S1—C9	104.87 (6)	C16—C15—H15A	108.6
C24—O4—C25	114.90 (11)	C1—C15—H15B	108.6
C1—N1—C8	108.36 (10)	C16—C15—H15B	108.6
C1—N1—S1	127.73 (9)	H15A—C15—H15B	107.5
C8—N1—S1	123.44 (9)	C21—C16—C17	118.73 (13)
C2—C1—N1	108.71 (11)	C21—C16—C15	122.10 (12)
C2—C1—C15	126.67 (11)	C17—C16—C15	119.15 (12)
N1—C1—C15	124.62 (11)	C18—C17—C16	120.63 (15)
C1—C2—C22	122.85 (12)	C18—C17—H17	119.7
C1—C2—C3	107.92 (11)	C16—C17—H17	119.7
C22—C2—C3	129.23 (12)	C19—C18—C17	120.32 (14)
C4—C3—C8	119.31 (11)	C19—C18—H18	119.8
C4—C3—C2	133.18 (12)	C17—C18—H18	119.8
C8—C3—C2	107.48 (11)	C18—C19—C20	119.64 (14)
C5—C4—C3	118.96 (12)	C18—C19—H19	120.2
C5—C4—H4	120.5	C20—C19—H19	120.2
C3—C4—H4	120.5	C19—C20—C21	119.97 (15)
C4—C5—C6	120.81 (12)	C19—C20—H20	120.0
C4—C5—H5	119.6	C21—C20—H20	120.0
C6—C5—H5	119.6	C16—C21—C20	120.67 (14)
C7—C6—C5	121.38 (12)	C16—C21—H21	119.7
C7—C6—H6	119.3	C20—C21—H21	119.7
C5—C6—H6	119.3	C23—C22—C2	128.59 (12)
C6—C7—C8	117.60 (12)	C23—C22—H22	115.7
C6—C7—H7	121.2	C2—C22—H22	115.7
C8—C7—H7	121.2	C22—C23—C24	119.08 (12)
C7—C8—C3	121.90 (12)	C22—C23—H23	120.5
C7—C8—N1	130.59 (12)	C24—C23—H23	120.5
C3—C8—N1	107.51 (11)	O3—C24—O4	122.91 (12)
C10—C9—C14	121.43 (12)	O3—C24—C23	125.98 (13)



C10—C9—S1	119.31 (10)	O4—C24—C23	111.10 (11)
C14—C9—S1	119.26 (10)	O4—C25—C26	107.16 (12)
C11—C10—C9	118.86 (13)	O4—C25—H25A	110.3
C11—C10—H10	120.6	C26—C25—H25A	110.3
C9—C10—H10	120.6	O4—C25—H25B	110.3
C12—C11—C10	120.13 (13)	C26—C25—H25B	110.3
C12—C11—H11	119.9	H25A—C25—H25B	108.5
C10—C11—H11	119.9	C25—C26—H26A	109.5
C11—C12—C13	120.56 (13)	C25—C26—H26B	109.5
C11—C12—H12	119.7	H26A—C26—H26B	109.5
C13—C12—H12	119.7	C25—C26—H26C	109.5
C14—C13—C12	120.27 (13)	H26A—C26—H26C	109.5
C14—C13—H13	119.9	H26B—C26—H26C	109.5
C12—C13—H13	119.9		
O1—S1—N1—C1	22.55 (12)	O2—S1—C9—C10	-148.51 (11)
O2—S1—N1—C1	152.10 (11)	N1—S1—C9—C10	98.91 (11)
C9—S1—N1—C1	-93.33 (11)	O1—S1—C9—C14	164.62 (10)
O1—S1—N1—C8	-166.17 (10)	O2—S1—C9—C14	31.32 (12)
O2—S1—N1—C8	-36.63 (11)	N1—S1—C9—C14	-81.26 (11)
C9—S1—N1—C8	77.95 (11)	C14—C9—C10—C11	1.3 (2)
C8—N1—C1—C2	1.03 (14)	S1—C9—C10—C11	-178.86 (11)
S1—N1—C1—C2	173.37 (9)	C9—C10—C11—C12	-0.2 (2)
C8—N1—C1—C15	-178.99 (11)	C10—C11—C12—C13	-1.2 (2)
S1—N1—C1—C15	-6.66 (18)	C11—C12—C13—C14	1.6 (2)
N1—C1—C2—C22	178.70 (11)	C12—C13—C14—C9	-0.5 (2)
C15—C1—C2—C22	-1.3 (2)	C10—C9—C14—C13	-1.0 (2)
N1—C1—C2—C3	-1.33 (14)	S1—C9—C14—C13	179.21 (11)
C15—C1—C2—C3	178.70 (12)	C2—C1—C15—C16	-98.29 (15)
C1—C2—C3—C4	-176.70 (13)	N1—C1—C15—C16	81.74 (15)
C22—C2—C3—C4	3.3 (2)	C1—C15—C16—C21	24.17 (18)
C1—C2—C3—C8	1.14 (14)	C1—C15—C16—C17	-157.46 (12)
C22—C2—C3—C8	-178.89 (12)	C21—C16—C17—C18	-1.6 (2)
C8—C3—C4—C5	0.24 (18)	C15—C16—C17—C18	179.96 (13)
C2—C3—C4—C5	177.87 (13)	C16—C17—C18—C19	0.0 (2)
C3—C4—C5—C6	1.37 (19)	C17—C18—C19—C20	1.1 (2)
C4—C5—C6—C7	-1.5 (2)	C18—C19—C20—C21	-0.5 (2)
C5—C6—C7—C8	0.0 (2)	C17—C16—C21—C20	2.2 (2)
C6—C7—C8—C3	1.65 (19)	C15—C16—C21—C20	-179.43 (13)
C6—C7—C8—N1	-177.70 (12)	C19—C20—C21—C16	-1.1 (2)
C4—C3—C8—C7	-1.79 (18)	C1—C2—C22—C23	-176.57 (13)
C2—C3—C8—C7	-179.98 (11)	C3—C2—C22—C23	3.5 (2)
C4—C3—C8—N1	177.69 (11)	C2—C22—C23—C24	177.74 (13)
C2—C3—C8—N1	-0.50 (13)	C25—O4—C24—O3	-0.3 (2)
C1—N1—C8—C7	179.12 (13)	C25—O4—C24—C23	179.10 (12)
S1—N1—C8—C7	6.38 (19)	C22—C23—C24—O3	1.0 (2)
C1—N1—C8—C3	-0.31 (13)	C22—C23—C24—O4	-178.39 (12)
S1—N1—C8—C3	-173.04 (8)	C24—O4—C25—C26	-171.93 (13)
O1—S1—C9—C10	-15.21 (13)		

## supplementary materials

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### Hydrogen-bond geometry (Å, °)

<i>D</i> —H··· <i>A</i>	<i>D</i> —H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> —H··· <i>A</i>
C7—H7···O2	0.95	2.36	2.9424 (17)	119
C15—H15A···O1	0.99	2.24	2.8474 (17)	118
C18—H18···O3 <sup>i</sup>	0.95	2.59	3.282 (2)	130
C10—H10···Cg1	0.95	2.75	3.3524 (14)	122
C11—H11···Cg2 <sup>ii</sup>	0.95	2.81	3.6263 (15)	145

Symmetry codes: (i)  $-x, -y+1, -z$ ; (ii)  $-x, y-1/2, -z+1/2$ .

Fig. 1

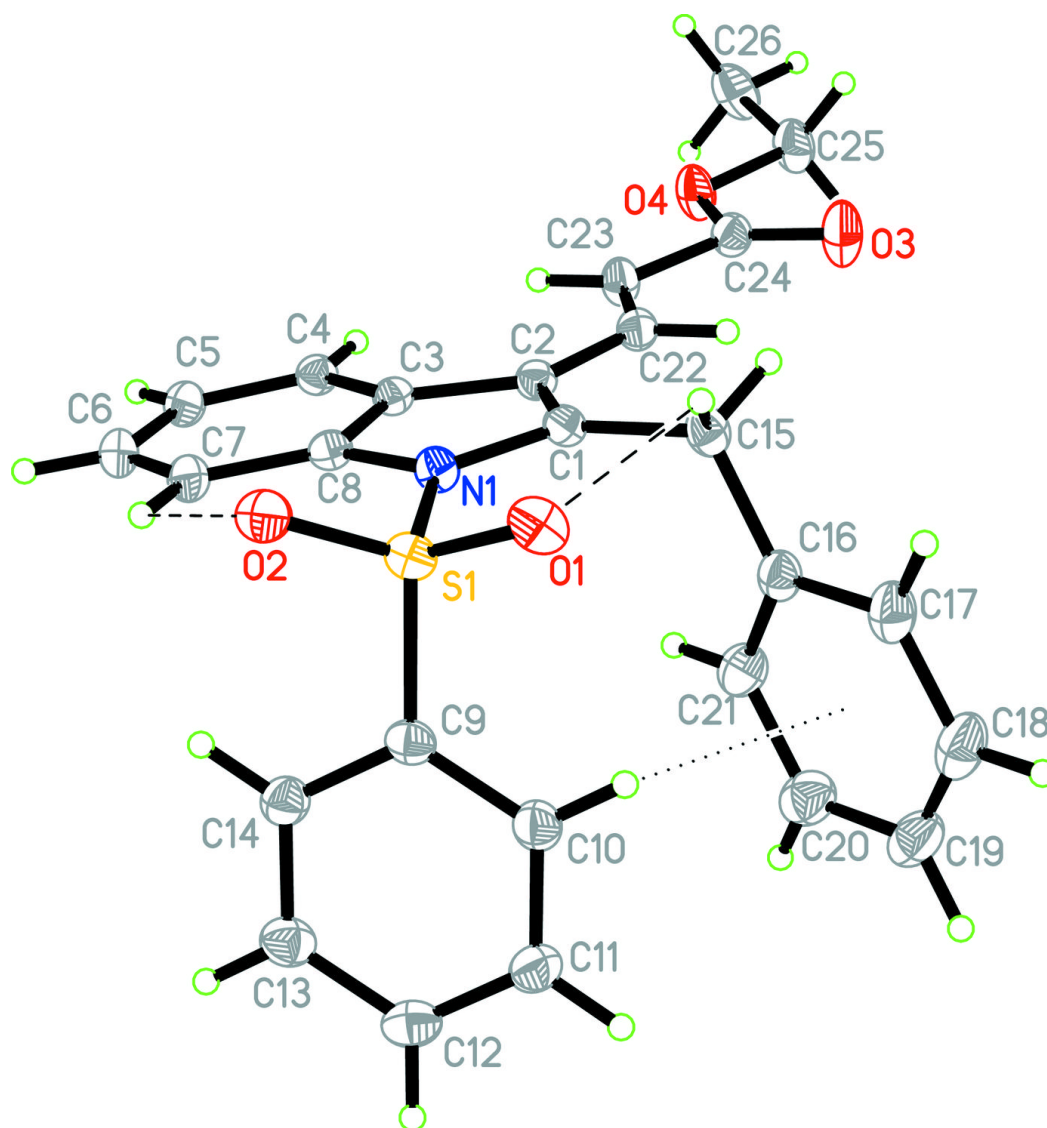


Fig. 2

